As a rule, the results are slightly higher by the Moore method than by the official method; this is only noticeably so in case of sample 1763. From this it would appear that the flocculent precipitate of iron and lime phosphate has but little to do with the low results usually obtained. It is to be noticed that only when the acid alcohol is used may the impurities contained in the platinum precipitate be neglected. Where plain alcohol is used it is usually necessary to dissolve the potassium chlorplatinate, wash, and reweigh the crucible to obtain the true weight of the potash.

It is obvious that after igniting the material it is not sufficient to treat it with water alone in order to dissolve the potash; such procedure does not always secure all the potash, some of it remaining undissolved, possibly in complex silicates or phosphosilicates, adhering to the dish. It was found that the potash so held could usually be readily recovered by dissolving the ignited material in the dish in dilute hydrochloric acid. This is not always true, however, the potash so held being dissolved but slowly in some instances, and while it appears that all potash may be recovered by prolonged treatment, I do not consider that this method has any advantages over the official method for regular fertilizer work, and further study was abandoned.

LEATHER AND PAPER LABORATORY.

NOTE.

An Explanation.—In the paper in the American *Geologist* on "The Dolomites of Eastern Iowa," which is abstracted in this Journal, **26**, R 373, one fact should have been stated. The author had previously determined that if the silica in these dolomites is in the form of a silicate, it is of a class that can be decomposed by digestion and evaporation with hydrochloric acid. This was ascertained by obtaining concordant results when treating the dolomite powder with a "fusion mixture" of sodium-potassium carbonate, and again by digesting with hydrochloric acid, that is, by evaporating to dryness, stirring until a fine, dry powder results. The residue in each case was treated with pure hydrofluoric and sulphuric acids. Also the residues obtained by dissolving the powder with hydrochloric acid on the one hand, and from digesting with hydrochloric acid on the other, were shown to be silica by the sulphuric-hydrofluoric method. The fusion process was, therefore, abandoned, partly on account of the difficulty of getting sodium and potassium carbonates free from silica. There is usually only a small amount of silica present in these rocks, often from 0.12 per cent. to 0.75 per cent. The paper referred to further shows that the evaporation to dryness and stirring to a fine powder is not necessary in these analyses, as the same result in silica can be obtained in two or three hours less time by simply dissolving the fine dolomite powder with hydrochloric acid, and considering the undissolved residue as silica, which it really is. This experiment, coupled with the work of the petrological microscope, showed that the silica present is not in the form of a silicate, but in the uncombined condition. The writer has worked with dolomites from various localities, but does not recall an instance where the silicate could not be decomposed by the digestion with hydrochloric acid. This, of course, would not be the case if the silicate consisted of particles of granite, syenite or similar rock.

NICHOLAS KNIGHT.

CORNELL COLLEGE,
October 12, 1904.

REVIEW.

RECENT WORK IN INORGANIC CHEMISTRY.

BY JAS. LEWIS HOWE.

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THE past twelve months have been characterized by an unusually large number of papers in the various journals in the field of inorganic chemistry. These papers are increasingly concerned with the application of the methods and theories of modern physical chemistry to the solution of the problems of inorganic chemistry, or perhaps we should more correctly say that the so-called physical chemistry is being more and more incorporated as an integral part of general chemistry. The result is an almost new inorganic chemistry. Another division of the subject which is beginning to yield important results is that of mixed organic-inorganic compounds, such as "complex" salts with organic acids. The rapid breaking-down of the lines between the different divisions of chemistry progresses, a growing protest against the breaking-up of our chemical societies into those confined to special departments.

During the past year the subject of greatest interest has been.